Topic X Test
Answer Section

MULTIPLE CHOICE

1. **ANS:** C  **PTS:** 1  **DIF:** L2  **REF:** p. 389
   **OBJ:** 12.1.2 Describe the quantities you can use to interpret a balanced chemical equation.
   **STA:** SC.912.P.8.9  **BLM:** application

2. **ANS:** B  **PTS:** 1  **DIF:** L1  **REF:** p. 401
   **OBJ:** 12.3.1 Explain how the amount of product in a reaction is affected by an insufficient quantity of any of the reactants.
   **STA:** SC.912.P.8.9 | SC.912.N.1.1  **BLM:** comprehension

3. **ANS:** B  **PTS:** 1  **DIF:** L3  **REF:** p. 401 | p. 402
   **OBJ:** 12.3.1 Explain how the amount of product in a reaction is affected by an insufficient quantity of any of the reactants.
   **STA:** SC.912.P.8.9 | SC.912.N.1.1  **BLM:** analysis

4. **ANS:** D  **PTS:** 1  **DIF:** L2  **REF:** p. 405
   **OBJ:** 12.3.2 Explain what the percent yield of a reaction measures.
   **STA:** SC.912.P.8.9 | SC.912.N.1.1  **BLM:** comprehension

5. **ANS:** C  **PTS:** 1  **DIF:** L1  **REF:** p. 405
   **OBJ:** 12.3.2 Explain what the percent yield of a reaction measures.
   **STA:** SC.912.P.8.9 | SC.912.N.1.1  **BLM:** knowledge

6. **ANS:** A  **PTS:** 1  **DIF:** L1  **REF:** p. 403
   **OBJ:** 12.3.1 Explain how the amount of product in a reaction is affected by an insufficient quantity of any of the reactants.
   **STA:** SC.912.P.8.9 | SC.912.N.1.1  **BLM:** comprehension

7. **ANS:** C  **PTS:** 1  **DIF:** L1  **REF:** p. 389
   **OBJ:** 12.1.2 Describe the quantities you can use to interpret a balanced chemical equation.
   **STA:** SC.912.P.8.9  **BLM:** knowledge

8. **ANS:** A  **PTS:** 1  **DIF:** L2  **REF:** p. 386 | p. 387
   **OBJ:** 12.1.2 Describe the quantities you can use to interpret a balanced chemical equation.
   **STA:** SC.912.P.8.9  **BLM:** application

9. **ANS:** D  **PTS:** 1  **DIF:** L2  **REF:** p. 389
   **OBJ:** 12.1.2 Describe the quantities you can use to interpret a balanced chemical equation.
   **STA:** SC.912.P.8.9  **BLM:** application

10. **ANS:** C  **PTS:** 1  **DIF:** L1  **REF:** p. 389
    **OBJ:** 12.1.2 Describe the quantities you can use to interpret a balanced chemical equation.
    **STA:** SC.912.P.8.9  **BLM:** knowledge

11. **ANS:** B  **PTS:** 1  **DIF:** L1  **REF:** p. 394
    **OBJ:** 12.2.2 Explain the general procedure for solving a stoichiometric problem.
    **STA:** SC.912.P.8.9 | SC.912.N.1.1  **BLM:** knowledge

12. **ANS:** C  **PTS:** 1  **DIF:** L1  **REF:** p. 389
    **OBJ:** 12.1.2 Describe the quantities you can use to interpret a balanced chemical equation.
    **STA:** SC.912.P.8.9  **BLM:** knowledge

13. **ANS:** B  **PTS:** 1  **DIF:** L2  **REF:** p. 386 | p. 387
    **OBJ:** 12.1.2 Describe the quantities you can use to interpret a balanced chemical equation.
    **STA:** SC.912.P.8.9  **BLM:** application

14. **ANS:** C  **PTS:** 1  **DIF:** L2  **REF:** p. 390
    **OBJ:** 12.2.1 Explain how mole ratios are used in chemical calculations.
    **STA:** SC.912.P.8.9 | SC.912.N.1.1  **BLM:** application

15. **ANS:** D  **PTS:** 1  **DIF:** L2  **REF:** p. 405
12.3.2 Explain what the percent yield of a reaction measures.

16. ANS: D  PTS: 1  DIF: L2  REF: p. 396

12.2.2 Explain the general procedure for solving a stoichiometric problem.

17. ANS: A  PTS: 1  DIF: L1  REF: p. 386

12.1.1 Describe how chemists use balanced chemical equations.

18. ANS: C  PTS: 1  DIF: L2  REF: p. 384